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How to find outer electron configuration of an element

The commonly used long form of the periodic table is designed to emphasize electron configurations. Since it is the outermost (valence) electrons which are primarily involved in chemical interactions between atoms, the last electron added to an atom in the building-up process is of far more interest to a chemist than the first. This last electron is called the distinguishing electron because it distinguishes an atom from the one immediately preceding it in the periodic table. The type of subshell (s, p, d, f) into which the distinguishing electron is placed is very closely related to the chemical behavior of an element and gives rise to the classification shown by the color-coding on the periodic table seen here. The representative elements are those in which the distinguishing electron enters an s or p subshell. Most of the elements whose chemistry and valence we have discussed so far fall into this category. Many of the chemical properties of the representative elements can be explained on the basis of Lewis diagrams. That is, the valences of the representative elements may be predicted on the basis of the number of valence electrons they have, or from the number of electrons that would have to be added in order to attain the same electron configuration as an atom of a noble gas. For representative elements the number of valence electrons is the same as the periodic group number, and the number needed to match the next noble-gas configuration is 8 minus the group number. This agrees with the valence rules derived from the periodic table, and results in formulas for chlorides of the first dozen elements that show the periodic variation of valence.
Element Atomic Weight Hydrogen Compounds Oxygen Compounds Chlorine Compounds Hydrogen 1.01 H2 H2O, H2O2 HCl Helium 4.00 None formed None formed None formed Lithium 6.94 LiH Li2O, Li2O2 LiCl Beryllium 9.01 BeH2 BeO BeCl2 Boron 10.81 B2H6 B2O3 BCl3 Carbon 12.01 CH4, C2H6, C3H8 CO2, CO, C2O3 CCl4, C2Cl6 Nitrogen 14.01 NH3, N2H4, HN3 N2O, NO, NO2, N2O5 NCl3 Oxygen 16.00 H2O, H2O2 O2, O3

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